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Molecular weight of
the gas = 29 ; $\gamma = 1.4$.

$1.4 \gamma - 1 = 1.4 - 1 = 0.4$

$1 \gamma + 1 = 1 + 1.4 = 2.4$

$\rho = 0.528 \text{ kg/m}^3$, $p = 430 \text{ kPa}$

$0.528 = 253.44 \text{ kPa}$

Fig. 11121 $T_1 = 303 \text{ K}$, $T_2 = 573.3 \text{ K}$

$\frac{p_2}{p_1} = \left(\frac{T_2}{T_1}\right)^{\frac{\gamma}{\gamma-1}} = \left(\frac{573.3}{303}\right)^{\frac{1.4}{0.4}}$

$= 303 \times 0.3333 = 101 \text{ kPa}$

573.3 K $R = 9 \text{ J/kg} \cdot \text{K}$, $\gamma = 1.4$

$= 0.2357 \text{ kJ/kg} \cdot \text{K}$

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$\frac{1}{1.0034} \text{ kJ/ kg}^{\circ}\text{C}$

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